

## Chapter 14 Acids And Bases Nhwweb Net

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Acids and Bases: Chapter 14 & 15

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A.P. Chemistry Practice Test: Ch. 14, Acids and Bases

CHAPTER 14 ACIDS AND BASES 347-log K<sub>a</sub> - log K<sub>b</sub> = -log K<sub>w</sub>, pK<sub>a</sub> + pK<sub>b</sub> = pK<sub>w</sub> = 14.00 (at 25°C) Exercises Nature of Acids and Bases 29. a. HCIO<sub>4</sub> (aq) + H<sub>2</sub>O(l) ⇌ H<sub>3</sub>O<sup>+</sup>(aq) + ClO<sub>4</sub><sup>-</sup>(aq) Only the forward reaction is indicated since HCIO<sub>4</sub> is a strong acid and is basically 100% dissociated in water. For acids, the dissociation reaction

Chapter 14 - Acids and Bases

A.P. Chemistry Practice Test: Ch. 14, Acids and Bases Name\_\_\_\_\_ MULTIPLE CHOICE. Choose the one alternative that best completes the statement or answers the question.

Chapter 14 - Acids and Bases - Mrs. Duffey - FHN

Major topics: types of acids, amphoterism, pH scale, simple pH calculations, strong acid calculations, weak acid calculations (ICE tables), & acid mixture problems.

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Chapter 14: Acids and Bases

CHAPTER 14 ACIDS AND BASES 5 When H<sub>3</sub>PO<sub>4</sub> is added to water, the three acids that are present are H<sub>3</sub>PO<sub>4</sub>, H<sub>2</sub>PO<sub>4</sub><sup>-</sup>, and HPO<sub>4</sub><sup>2-</sup>. H<sub>3</sub>PO<sub>4</sub>, with the largest K<sub>a</sub> value, is the strongest of these weak acids. The conjugate bases of the three acids are H<sub>2</sub>PO<sub>4</sub><sup>-</sup>, HPO<sub>4</sub><sup>2-</sup>

14.1 Brønsted-Lowry Acids and Bases – Chemistry

A vodcast I made for watching later in the chapter (after we have talked about strong and weak acids and bases in class as well as salts). It goes over all the different categories of pH problems in this chapter (strong acid, weak base, salt of weak acid/strong base, etc) and how to solve them.

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Arrhenius Acid-Base Theory Arrhenius Acid -a hydrogen containing compound that ionizes to produce hydrogen ions ( H<sup>+</sup>) when dissolved in water Because molecular acids are not made of ions, they cannot dissociate. HCl (aq) ⇌ H<sup>+</sup>(aq) + Cl<sup>-</sup>(aq) They must be pulled apart, or ionized, by the water. Chapter 14: Acids and Bases Ch 14 Page 1

Chapter 14: Acids and Bases

Acids and Bases: Chapter 14 & 15 . HW: •Read Ch 14: •Fill in as much of the acid base table as you can, as you read . Acid base conductivity and reactivity Conductivity, Reactivity,, Hydrochloric\*acid\* high high Acidic\*acid\* \*\* low\* medium Dis4led\*water\* none\* none\* Ammoniumhydroxide med\* none\* Sodiumhydroxide\* high none\* \*

CHAPTER FOURTEEN ACIDS AND BASES

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CHAPTER 14 REVIEW . Acids and Bases. SHORT ANSWER Answer the following questions in the space provided. 1. a. Write the two equations that show the two-stage ionization of sulfurous acid in water. stage 1: H<sub>2</sub>SO<sub>3</sub>(aq) + H<sub>2</sub>O(l) ⇌ H<sub>3</sub>O<sup>+</sup>(aq) + HSO<sub>3</sub><sup>-</sup>(aq) b.

14 Acids and Bases

Acids and Bases Know the definition of Arrhenius, Bronsted-Lowry, and Lewis acid and base. Autoionization of Water Since we will be dealing with aqueous acid and base solution, first we must examine the behavior of water.

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Chapter 14 - Acids and Bases . 14.1 The Nature of Acids and Bases . A. Arrhenius Model 1. Acids produce hydrogen ions in aqueous solutions 2. Bases produce hydroxide ions in aqueous solutions B. Bronsted-Lowry Model 1. Acids are proton donors 2. Bases are proton acceptors 3. H<sub>3</sub>O<sup>+</sup> is called the hydronium ion C. Conjugate Acid- Base Pairs 1.

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CHAPTER 14 REVIEW Acids and Bases SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Answer the following questions according to the Brønsted-Lowry definitions of acids and bases: HSO<sub>3</sub><sup>-</sup> a. What is the conjugate base of H<sub>2</sub>SO<sub>3</sub>? NH<sub>3</sub> b. What is the conjugate base of NH<sub>4</sub><sup>+</sup>?

Chapter 14 - Acids and Bases — HCC Learning Web

Chapter 14. Acid-Base Equilibria. 14.1 Brønsted-Lowry Acids and Bases Learning Objectives. By the end of this section, you will be able to: Identify acids, bases, and conjugate acid-base pairs according to the Brønsted-Lowry definition; Write equations for acid and base ionization reactions;

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