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*understood,
achievement does
not suggest that
you have wonderful
points.*

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answer sheet!!

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choices to answer
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E.C. The
approximate
composition of
naturally occurring
magnesium is as
follows: 79%
magnesium-24,
10%
magnesium-25,
and 11%*

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*magnesium-26.
Calculate the
atomic mass of
magnesium.*

*Organic Chemistry
I*

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Review. What are
the symbols for the
elements: lead,
silver, gold,
potassium,*

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mercury, iron, and sodium. What are the SI units for: mass, length, volume, energy, time? The density of aluminum is 2.7 g/cm^3 . A piece of aluminum that occupies a volume of 21.4 cm^3 would have a mass of?

Chapter 4

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Atoms of different
elements can
physically combine
in simple whole-
number ratios to
form compounds 4.
Chemical reactions
occur when atoms
are separated,
joined, or
rearranged. Atoms
of one element,*

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however, are never changed into atoms of another element as a result of a chemical reaction.

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Chapters 4-6 -You
may use a periodic
table and*

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calculator for this test. -Please show all work and calculations.

-Please answer all questions on a separate sheet of paper.

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Test With Answer

for Chemistry

chapter 4

(Chemical Bonding)

In this

topic, Student

should be able to :

a) Characterize

electrovalent

(ionic) bond as in

sodium chloride

and calcium oxide.

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Gravity. Created

by. Hoolia. Terms

in this set (90)

Democritus.

ancient greek

thinker believed

that matter could

be divided into

smaller and smaller

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*particles until a
basic particle was
reached. He was
the first to suggest
the existence of
atoms*

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ORGANIC
CHEMISTRY I CHEM
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C. Biewer biewerm
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Acid | Solution

Chapter 4 -

Reactions in

Aqueous Solution:

Part 1 of 6 ... I'll

continue our

General Chemistry

course by teaching

you the meaning of

the terms solvent,

solute, electrolyte,

and nonelectrolyte

...

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*Name _____ SHORT
ANSWER. Write the
word or phrase that
best completes
each statement or
answers the*

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question. 1)What is the solvent in an aqueous solution?
ESSAY. Write your answer in the space provided or on a separate sheet of paper.
2)Explain why water is a good solvent for ionic substances. ...

FSC Part 1

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Chemistry Chapter
4 Liquids and
Solids Online ...

4. Determine the
volume of 0.100M
 Na_3PO_4 required to
precipitate all the
lead (II) ions from
150.0 mL of
0.250M $\text{Pb}(\text{NO}_3)_2$.
(4 points) 5.

Determine the
mass of barium
sulfate that is

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*produced when
100.0 mL of 0.100
M BaCl₂ is mixed
with 100.0 mL of
0.100 M
Fe₂(SO₄)₃. (3
points) 6.*

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4 Practice Multiple
Choice Questions

Multiple Choice

Identify the choice
that best

completes the
statement or
answers the

question. ____ 1.

What mass of silver
nitrate, AgNO_3 , is

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*required to prepare
800. g of 3.50%
solution of AgNO₃?*
a. 24.6 g b. 26.7 g
c. 27.0 g d. 25.5 g

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