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The maximum number of electrons an orbital can contain is 2 electrons. Describe the relative orientations of the orbitals related to an atom's 2p sublevel. The relative orientations of the orbitals in an atom's 2p sublevel is 3 orbitals, 2px, 2py and 2pz that look like figure eights and are perpendicular to each other

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states that a maximum of two electrons can occupy a single atomic orbital, but if

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only if the electrons have opposite spins states that each electron occupies the lowest energy orbital available states that single electrons electrons with the same spin must occupy each-energy orbital before additional electrons with opposite spins can occupy the ...

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A three-dimensional region around the nucleus of an atom that.... electrons occupy

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orbitals of the lowest energy levels first. an arrangement of electrons around the nucleus of an atom (loo.... The amount of energy an electron can have.

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