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3 Atoms: The Building Blocks of Matter - Ms. Sheehan's ...

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2 Measurements and Calculations

CHAPTER 10 REVIEW States of Matter SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Match description on the right to the correct crystal type on the left. b ionic crystal (a) has mobile electrons in the crystal c covalent molecular crystal (b) is hard, brittle, and nonconducting a metallic crystal (c) typically has the lowest melting point of the four

3 Atoms: The Building Blocks of Matter

CHAPTER 3 REVIEW Atoms: The Building Blocks of Matter SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. In cathode-ray tubes, the cathode ray is emitted from the negative electrode, which is called the cathode . 2. The smallest unit of an element that can exist either alone or in molecules containing the

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CHAPTER 3 REVIEW Atoms: The Building Blocks of Matter SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. In cathode-ray tubes, the cathode ray is emitted from the negative electrode, which is called the cathode . 2. The smallest unit of an element that can exist either alone or in molecules containing the

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Modern Chemistry Chapter 2-3 Review Answers

SECTION 1 continued 6. Some binary compounds are ionic, others are covalent. The type of bond favored partially depends on the position of the elements in the periodic table.

5 The Periodic Law

CH 1 Reading Assignment Modern Chemistry. CH 1 Vocabulary-New. CH 1 Mixed Questions. CH 1 Matter & Energy Vocabulary-New ... Chapter 3 CH 3 Atomic Structure. Reference: CH 3 Atomic Structure Organizer. ... Ch 6 Section Review 6.3. Ch 6 Section Review 6.4. Ch 6 Section Review 6.5.

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Chapter Tests With Answer Key Modern Chemistry, 2006 ...

MODERN CHEMISTRY The atomic number of nickel-60 is 28. How many neutrons does this iso- tope have? Carbon-14 has 8 neutrons. What is the atomic number of carbon-14? 11 CHAPTER 3 TEST HR W materiai copyrighted under notice appeafing earlier in this work.

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3:3 mol H₂; or their reciprocals 7. Given the following equation: 4NH₃(g) + 6NO(g) → 5N₂(g) + 6H₂O(g) 1 mol NO:1 mol H₂O a. What is the mole ratio of

NO to H₂O? 3 mol NO:2 mol NH₃ b. What is the mole ratio of NO to NH₃? 0.360 mol c. If 0.240 mol of NH₃ react according to the above equation, how many moles of NO will be consumed? 8.

7 Chemical Formulas and Chemical Compounds

CHAPTER 5 REVIEW The Periodic Law SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. c In the modern periodic table, elements are ordered (a) according to decreasing atomic mass. (b) according to Mendeleev's original design. ... 40 THE PERIODIC LAW MODERN CHEMISTRY ...

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CHAPTER 6 REVIEW Chemical Bonding SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. a The notation for sodium chloride, NaCl, stands for one (a) formula unit. (c) crystal. (b) molecule. (d) atom. 2. d In a crystal of an ionic compound, each cation is surrounded by a number of (a) molecules. (c) dipoles. (b) ...

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6 Chemical Bonding

3. Compare and contrast a model with a theory. Theories are broad generalizations used to explain observations. Models are a physical object used to illustrate or explain complex concepts or an explanation of how phenomena occur and how data and events are related. MODERN CHEMISTRY MEASUREMENTS AND CALCULATIONS 9

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10 States of Matter - Ms. Agostine's Chemistry Page

The Atoms: The Building Blocks of Matter chapter of this Holt McDougal Modern Chemistry Companion Course helps students learn the essential lessons associated with atoms.

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