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Chem 115 POGIL Worksheet - Week #7 - Answers First Law, Enthalpy, Calorimetry, and Hess's Law Key Questions and Exercises 1. What are the two major ways in which the internal energy of an object can be categorized? How do these ways differ from one another? An object can have kinetic energy and potential energy. The kinetic energy arises from

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Calorimetry - Measurement of Heat Energy

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If 4.168 kJ of heat is added to a calorimeter containing 75.40 g of water, the temperature of the water and the calorimeter increases from 24.58 degree C to 35.82 degree C. Calculate the heat capac...

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In this POGIL, you will explore how mass, temperature, heat energy, ... 45. Does it take more energy to heat water or ice? Explain your answer, but be careful. The water from the problem underwent a 100°C temperature change, but ice only increased by 50°C. -100-50 0 50 100 150

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In this activity you will explore how mass, temperature, heat energy, and the type of substance are related. a rate of 30 Joules per minute Calorimetry 1 Model 1 - A Pot of Water Before heating 24.5 °C Temp. 4 qt. saucepan 24.5 °C Temp. 8 qt. stockpot After heating for 5 minutes at

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a calorimetry calculation is a procedure that allows you to have more accurate answers since you are also considering the heat that is absorbed by the calorimeter and is thus not recorded. 2° HCl HNO, HNO, |24. student exploration calorimetry lab answers Media Publishing eBook, ePub, Kindle PDF View ID f4352a23f Apr 29, 2020 By Edgar Wallace before read online student exploration ...

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Calorimetry - Measurement of Heat Energy 7. Determine the mass of a water sample that is heated from an initial temperature of 25 0C to a final temperature of 100. 0C following the addition of 1200 J of heat energy. 8. A 100. g sample of pure lead is heated from 10.0oC to 197.5oC by the addition of 3000. J of heat energy.

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