

## The Ultimate Chemical Equations Handbook Answers Chapter 3

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Edition ...

The Ultimate Chemical Equations Handbook . Batavia, IL: Flinn Scientific, Inc. Correlating questions (ii) from: Rodney Schmitz (2007) Moore County Schools, NC For each of the following eight reactions, in part (i) write a BALANCED equation and in part (ii) answer the question about the reaction.

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The Ultimate Chemical Equations Greek, or German names. Ag Sn w (ferrum) (natrium) (argentum) (stannum) (plumbum) (wolfram) iron sodium silver tin lead tungsten (cuprum) (kalium) (hydrargyrum) (stibium) (aurium) copper potassium mercury antimony gold Symbols are used as a sort of shorthand in writing the names of elements.

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No reaction 8. Calcium metal is added to a solution of nitrous acid.  $\text{Ca(s)} + 2\text{HNO}_2(\text{aq}) \rightarrow \text{Ca(NO}_2)_2(\text{aq}) + \text{H}_2(\text{g})$  9. A pea-sized piece of lithium is added to water.  $2\text{Li(s)} + 2\text{H}_2\text{O(l)} \rightarrow 2\text{LiOH(aq)} + \text{H}_2(\text{g})$  10. A solution of iron(III) chloride is poured over a piece of platinum wire. No reaction

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tion. The equations are of mixed types. This Equations

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section is worth 15 points and is equal to 15 per. cent of the free-response grade. It is for this section of the exam that this handbook was written. All AP equations "work." This means that in every case, a reaction will occur—there are no "no reac-

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## Chapter 3

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The Ultimate Chemical Equations Handbook Student edition by George R. Hauge, Jane D. Smith Published June 2001 by Flinn Scientific. Written in English.

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Oxidation Numbers of Representative Element Cations and Anions

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Ultimate equation answers

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Cations. Come from metals that lose electrons in order to become isoelectronic with a noble gas. Anions. Come from nonmetals that gain electrons to become isoelectronic with a noble gas. Oxidation. The process of losing electrons to become isoelectronic with a noble gas.

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The Ultimate Chemical Equations Handbook

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The Ultimate Chemical Equations Handbook is new and revised, offering thorough, comprehensive examples and exercises that provide continuous reinforcement to improve students' chemical literacy skills.

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